| Name: | KEY |
|-------|-----|
| | |

21. (A) (C) (D)

22. (A) (B) (C)

23. (A) (B) (C)

24. (A) (B) (D)

25. (A) (B) (C)

Science 9 – Unit Test – Matter and Chemical Change

Version 1

Part A - Multiple Choice

- 1. **B**CD
- 2. **B**CD
- 3. **(A) (— (C) (D)**
- 4. (A) (C) (D)
- 5. (A) (B) (C)
- 6. **B**CD
- 7. (A) (B) (C)
- 8. **(A) (B) (D)**
- 9. (A) (B) (D)
- 10. **(A) (C) (D)**

- 11. (A) (B) (D)
- 12. (A) (B) (D)
- 13. **B**© **D**
- 14. (A) (B) (C) (
- 15. **(A) (B) (D)**
- 16. **B**© **D**
- 17. (A) (B) (D)
- 18. **B**CD
- 19. **A** C •
- 20. (A) (B) (D)

Numerical Response

1. 1 2 3

2. 1 1

:

1 2 . 0

Part B - Written Response

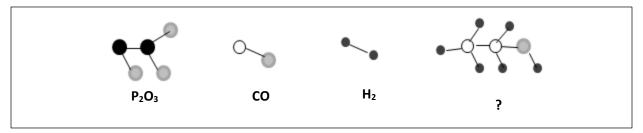
Written Response 1

Fill in the missing information from the table below (5 marks)

| Chemical Formula | Type of compound (ionic or molecular?) | Chemical name |
|-------------------------------|--|-------------------------|
| MgO | ionic | magnesium oxide |
| Na ₂ S | ionic | sodium sulfide |
| CH ₄ | molecular | carbon tetrahydride |
| P ₂ O ₅ | molecular | diphosphorus pentaoxide |
| Be₃N₂ | ionic | beryllium nitride |

Written Response 2

Use the following information to answer Written Response 2



a.) What is the **chemical formula** for the unknown molecule? (1 mark)

C₂H₆O

b.) Write the **chemical equation** for the combustion of the unknown molecule (Hint: the unknown molecule is the fuel source) (2 marks)

This question was removed from the test, but you need to know the general formulas for combustion and corrosion:

Combustion: $CH_4 + O_2 \rightarrow CO_2 + H_2O$

Written Response 3

Corrosion: Fe + $O_2 \rightarrow Fe_2O_3$

Use the following information to answer Written Response 3

Billy is trying to determine the chemical formula for the compound that forms between rubidium and sulfur. He does the following work:

Step 1: Rb²⁺ S²⁻ Step 4: <u>RbS</u>₂

Step 2: Rb1+ Answer: The compound that forms between rubidium and sulfur is

Step 3: 2+ 2- RbS2.

Billy's friend Alice says that he is wrong, and that the right answer is Rb₂S. Who is correct? **Explain your answer**. (2 marks)

Alice is right. Billy wrote steps 1-3 correctly, but he should have Rb₂S for Step 4, because two

Rubidium are required to balance the 2- of sulfur.