Science 9 - Unit Test - Matter and Chemical Change

## Version 1

Part A - Multiple Choice

| 1. (B) (C) (D) | 11. (A) (B) (D) | 21. ( $\overline{\text { a }}$ ( ( $)^{\text {( }}$ ( $)$ |
| :---: | :---: | :---: |
|  | 12. ( $\overline{\text { a }}$ ( $(\bar{B})$ (D) | 22. (A) $(\overline{\text { B }})(\overline{\text { C }})$ |
| 3. (A) (C) (D) | 13. (B) (C) (D) | 23. (A) $(\overline{\text { B }})(\overline{\text { C }})$ |
| 4. (A) (C) (D) | 14. (A) $(\overline{\text { B }})(\overline{\text { C }})$ | 24. (A) (B) ( D ${ }^{\text {( })}$ |
| 5. (A) $(\overline{\text { B }})(\overline{\mathrm{C}})$ | 15. ( $\overline{\text { A }})(\overline{\text { B }})(\overline{\text { D }})$ | 25. (A) $(\overline{\text { B }})(\overline{\text { C }})$ |
| 6. (B)(C) (D) | 16. (B) (C) (D) |  |
| 7. (A) $(\overline{\text { B }})(\overline{\text { C }})$ | 17. (A) (B) (D) |  |
| 8. (A) (B) (D) | 18. (B) (C) (D) |  |
| 9. (A) (B) (D) | 19. (-A) (C) (D) |  |
| 10. (A) (C) (D) | 20. (A) (B) (D) |  |

## Numerical Response

1. 

| 1 | 2 | 3 |  |
| :--- | :--- | :--- | :--- |

2. 


3.

| 1 | 2 |  |  |
| :--- | :--- | :--- | :--- |

## Part B - Written Response

## Written Response 1

Fill in the missing information from the table below (5 marks)

| Chemical Formula | Type of compound (ionic <br> or molecular?) | Chemical name |
| :---: | :---: | :---: |
| MgO | ionic | magnesium oxide |
| $\mathrm{Na}_{2} \mathrm{~S}$ | ionic | sodium sulfide |
| $\mathrm{CH}_{4}$ | molecular | carbon tetrahydride |
| $\mathrm{P}_{2} \mathrm{O}_{5}$ | molecular | diphosphorus pentaoxide |
| $\mathrm{Be}_{3} \mathrm{~N}_{2}$ | ionic | beryllium nitride |

## Written Response 2

Use the following information to answer Written Response 2

$\mathrm{P}_{2} \mathrm{O}_{3}$

CO

$\mathrm{H}_{2}$

?
a.) What is the chemical formula for the unknown molecule? (1 mark)
$\qquad$
$\mathrm{C}_{2} \mathrm{H}_{6} \mathrm{O}$
b.) Write the chemical equation for the combustion of the unknown molecule (Hint: the unknown molecule is the fuel source) (2 marks)

> This question was removed from the test, but you need to know the general formulas for combustion and corrosion:

Combustion: $\mathrm{CH}_{4}+\mathrm{O}_{2} \rightarrow \mathrm{CO}_{2}+\mathrm{H}_{2} \mathrm{O}$
Written Response 3 Corrosion: $\mathrm{Fe}+\mathrm{O}_{2} \rightarrow \mathrm{Fe}_{2} \mathrm{O}_{3}$
Use the following information to answer Written Response 3

Billy is trying to determine the chemical formula for the compound that forms between rubidium and sulfur. He does the following work:

| Step 1: RbIt | $S^{2-}$ | Step 4: $\mathrm{RbS}_{2}$ |
| :---: | :---: | :---: |
| Step 2: $\mathrm{Rb}^{1+}$ |  | Answer: The compound that forms |
|  |  | between rubidium and sulfur is |
| Step 3: $2+$ | $2-$ | $\mathrm{RbS}_{2}$. |

Billy's friend Alice says that he is wrong, and that the right answer is $R b_{2} S$. Who is correct? Explain your answer. (2 marks)

Alice is right. Billy wrote steps 1-3 correctly, but he should have $\mathrm{Rb}_{2} \mathrm{~S}$ for Step 4, because two
Rubidium are required to balance the 2 - of sulfur.

